

Chemistry Review



1. Everything in the universe is made of _____.
2. The measurement of the amount of matter in an object is called _____.
3. What are the 3 states of matter?
 - A. _____
 - B. _____
 - C. _____
4. Charged particles that move around an atom's nucleus are called _____.
5. Chemical bonds are broken, atoms are rearranged, and new bonds are formed during _____.
6. Atoms with filled outermost energy levels tend _____ to participate in chemical reactions.
7. A pure substance that cannot be broken down is called an _____.
8. The simplest part of an element is an _____.
9. The central core of an atom is called the _____.
10. In an ionic bond, _____ atoms of _____ charge are held together by _____ attraction.
11. The part of an atom that has a neutral charge is a _____.
12. Most of the mass of an atom is found in the _____.
13. A pure substance made up of atoms of one or more elements is called a _____.
14. Most atoms tend to undergo _____, combining in ways that cause their atoms to become more _____.
15. When two atoms share one or more electrons, it is called _____.

_____.

16. A bond formed by electrical attraction between two opposite charged ions is called _____.

17. The ability to do work or cause change is _____.

18. A redox reaction involves the _____ of _____ between atoms.

19. The amount of energy needed to start a chemical reaction is the reaction's _____.

20. A substance that neutralizes small amounts of acids or bases added to a solution is a _____.

21. A chemical reaction that can proceed forward or backward is a _____.

22. Sodium chloride (table salt) is an example of a compound formed by _____.

23. The positive charge part of an atom is called a _____.

24. A particle composed of one or more atoms is a _____.

25. Chemical reactions that release free energy are called _____.

26. Chemical reactions that absorb free energy are called _____.

27. The loss of one or more electrons is called _____.

28. The gaining of one or more electrons is called _____.

29. The breaking apart of water molecules into two ions of opposite charge is called _____.

30. An atom has six electrons, what is its atomic number? _____
Name? _____ It is a stable or unstable atom? _____.

31. Why is it necessary for oxidation and reduction reactions to occur in pairs?

32. What is the role of enzymes in chemical reactions occurring in living things?

33. Describe the relationship between the solute, the solvent, and the concentration of a solution.

34. List 2 characteristics of acids and 2 characteristics of bases.

35. What is the pH Scale, and what does its range of values mean?